Coloured Glasses

Colour is the light that transmits through or reflects off an object. The colour in glass comes from one, or a combination of factors or materials. Colour is rarely a constant produced by a particular oxide in isolation. The arrangement of electrons surrounding the atoms of a particular colouing material are affected by the light energy and by the magnetic energy of adjacent atoms. The remainder of the batch can have significant effect i.e. NiO in a Soda-Lime batch (or lead or barium) will create brown/grey colours – in a Potash Batch (potassium Carbonate) the nickel will create a violet glass.

The oxidising or reducing atmosphere (furnace conditions) can have a significant effect as some of the colours can exist in the glass in more than one valent state. Iron oxide can reduce to ferrous oxide and produce blue-green or it can oxidise to ferric oxide to produce yellow-greens and browns of bottle glass. Copper oxide is blue-green in oxidised alkaline glass but can be reduced to reds or even to produce a thin metallic layer on the glass surface. Glasses coloured by copper, gold, selenium and silver often need to be reheated to enable the colour to strike. The correct temperature and time required for them to strike varies considerably and will only be right for that furnace, or glory hole, at exactly those conditions, at that humidity, etc. It should be repeatable but over confidence beware. Selenium is a major problem when instead of red-orange a black, brown, green or yellow may appear due to its different valent states depending if oxidised or reduced.

Sometimes several "colouring oxides" are added to create an overall colour i.e. "black" is made up of oxides each affecting light transmission at different wavelengths:- blue, purple, grey, green, red etc. Sometimes two different materials are added which combine together i.e. Cadmium Sulphide and Selenium create Cadmium Selenide which creates a ruby red.

Colour	Colour Variation	Material (combination)	Conditions
Blues	Blue	CuO & CoO	
	Blue (Violet tint)	CoO	
	Sky Blue	CuO	
	Blue Green	Fe2O3 + CoO, CuO + Cr2O3, FeO	
Browns	Red to Yellow Brown	MnO2 + Fe2O3	Oxidised
	Carbon Amber	C + Na2SO4 + (FeS2)	Reduced
	Brown Orange	Se (red) + Fe2O3	Not Oxidised
	Brown / Grey	NiO	
Grey	Grey	MnO2 + Fe2O3 + CuO	Oxidised
	Grey	NiO + above in Soda-Lime glasses	
Green	Antique Green	Fe2O3 +(C + Cr2O3)	Reduced
	Emerald Green	Cr2O3	Reduced
	Grass Green	Cr2O3 + CuO	Oxidised
	Yellow Green	Cr2O3	Oxidised
		Pr6O11	
		K2Cr2O7	
		U3O8 (fluorescent)	Reduced

Green Blue FeO Redu Cr2O3 + CuO Redu Moss Green Cr2O3 + CoO Redu Cr2O3 + NiO2 Redu Olive Green Fe2O3 + CrO3 Oxid Cr2O3 + MnO2 Oxid	iced iced iced
Moss Green Cr2O3 + CoO Redu Cr2O3 + NiO2 Redu Olive Green Fe2O3 + CrO3 Oxid Cr2O3 + MnO2 Oxid	iced iced
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Olive Green Fe2O3 + CrO3 Oxid Cr2O3 + MnO2 Oxid	
Cr2O3 + MnO2 Oxid	
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Yellow Yellow / Green tint CdS Redu	ced
Intense Yellow U3O8 + Pr6O11 Oxid	ised
Gold Yellow Se + MnO2 in PbO glasses	
Canary Yellow TiO2 + Ce2O3	
Yellow / green fluorescence U3O8 Oxid	ised
Amber Yellow MnO2 + FeO Oxid	ised
Se in PbO glasses	
Orange Orange Red CdS + Se Sligh Redu	•
Orange Brown Se	
Rose Rose / Pink Se not in PbO glasses Redu	ced
Red Ruby Se + CdS not in PbO glasses Redu	ced + on ating
Au	
	ced + on ating
Red Violet MnO2 + Se Oxid	ised
Wine Red Nd2O3 + Se Not O	Oxidised
Black Black FeS	
Black Black FeS MnO2+Fe2O3+NiO+CoO+CuO Oxid	ised
MnO2+Fe2O3+NiO+CoO+CuO Oxid	iced
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